Writing and balancing equations, ionic equations And net ionic equations.	NameKEYBlock Date
From the following sentence write the balanced equation s identify the spectator ions.	, ionic equation and net ionic equation as well
 This reaction comes from the video we watched in Aqueous potassium iodide is mixed with a solution solution of potassium nitrate and a yellow solid le a. Write out the balanced equation. 	on of lead (II) nitrate. The result is an aqueous
2KI(ag) + 1Pb(NO)	$(0q) \rightarrow 2 KNO_3(aq) + 1PbI_2(s)$
b. Write out the ionic equation.	3
2K+cag> +2I-cag> + Pb2+(ag)	2NO_(aq) -2K+(aq) +2NO_(aq) + PbI_(
c. Write out the net ionic equation.	3 1 - 1 - 3 - 1 - 2
$2I^{-}(aq) + Pb^{2+}(aq)$	→ PhI.(s)
d. What are the spectator ions from the reaction	
K^+ and NO_2^-	
The Charles of the Children of	
 This reaction comes from your notes on page 241-2 Aqueous potassium chromate is mixed with a soluti aqueous solution of potassium nitrate and a yellow s a. Write out the balanced equation. 	on of barium nitrate. The result is an
K, CrO, cago + Ba(NO3),	(aq) - 2 KNO3(aq) + BaCrO2(s)
b. Write out the ionic equation.	
2K+(aq) + GrO2-(aq) + Ba2+(aq)	+ 2NO_caq) -2K+(aq) +2NO_caq) + Bac
c. Write out the net ionic equation.	
CrO_{q}^{2} -(aq) + Ba ²⁺ (aq) \rightarrow	BaCrO. (5)
d. What are the spectator ions from the reaction	

K+ and NO,-