

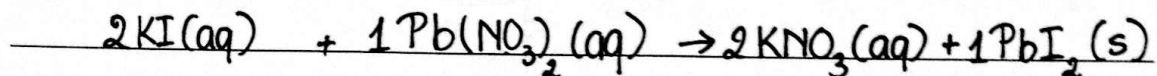
Writing and balancing equations, ionic equations
And net ionic equations.

Name KEY
Block _____ Date _____

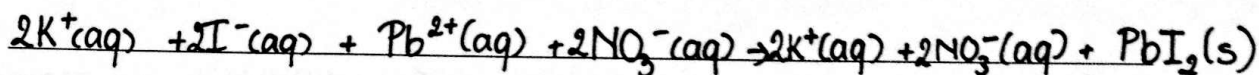
From the following sentence write the balanced equation, ionic equation and net ionic equation as well as identify the spectator ions.

1. This reaction comes from the video we watched in class and that is presented on our website. Aqueous potassium iodide is mixed with a solution of lead (II) nitrate. The result is an aqueous solution of potassium nitrate and a yellow solid lead (II) iodide precipitate.

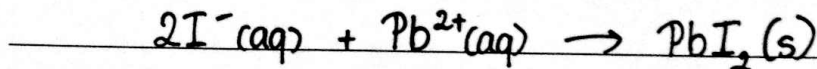
a. Write out the balanced equation.



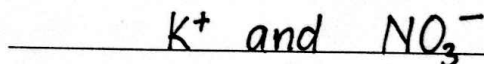
b. Write out the ionic equation.



c. Write out the net ionic equation.



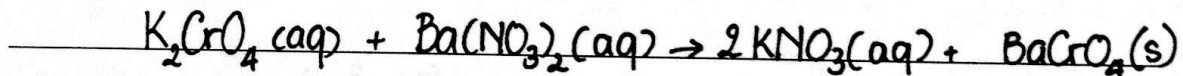
d. What are the spectator ions from the reaction?



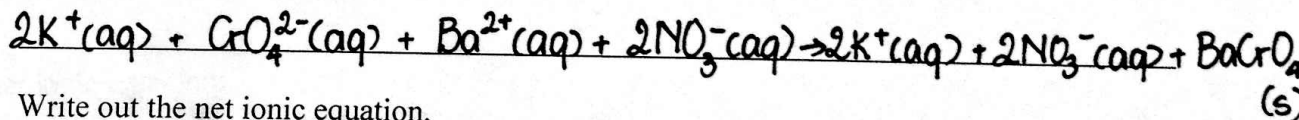
2. This reaction comes from your notes on page 241-244.

Aqueous potassium chromate is mixed with a solution of barium nitrate. The result is an aqueous solution of potassium nitrate and a yellow solid barium chromate precipitate.

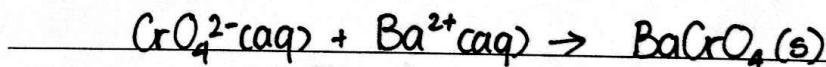
a. Write out the balanced equation.



b. Write out the ionic equation.



c. Write out the net ionic equation.



d. What are the spectator ions from the reaction?

