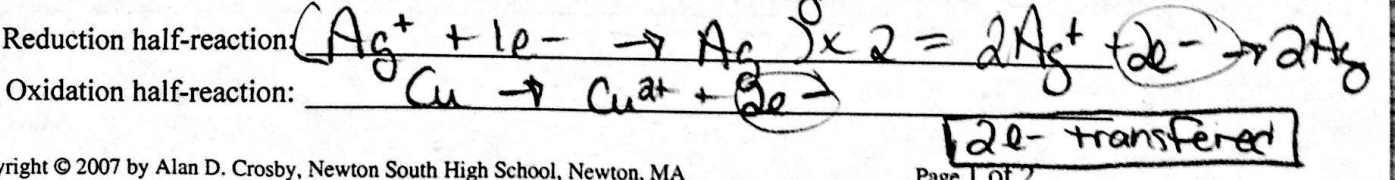
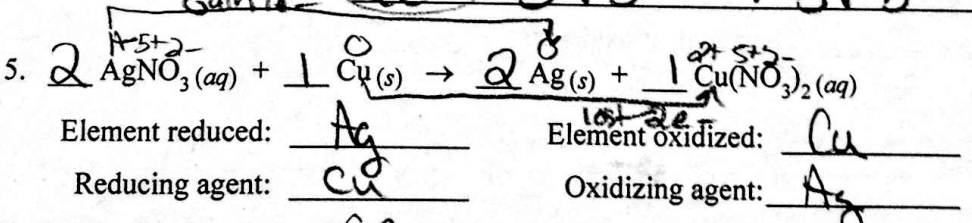
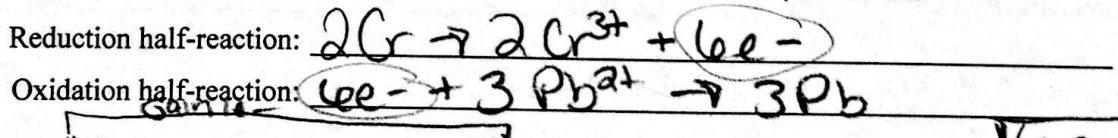
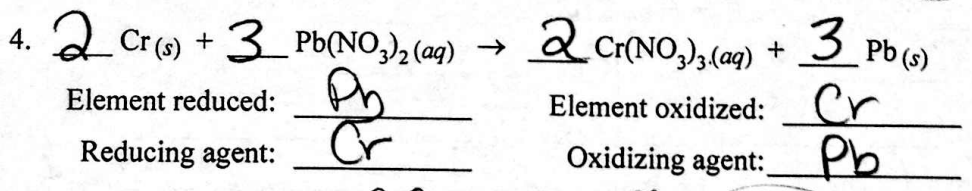
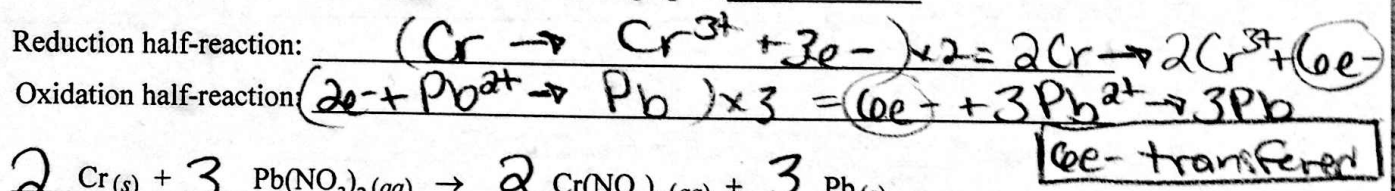
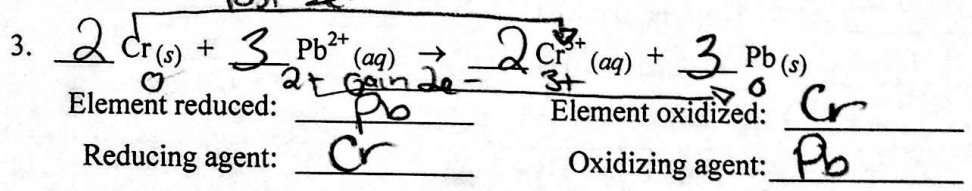
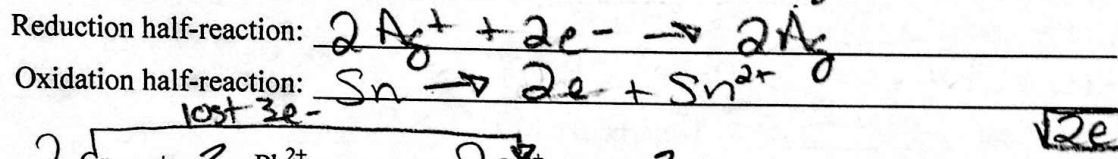
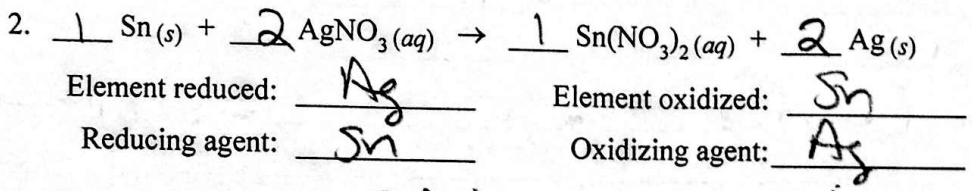
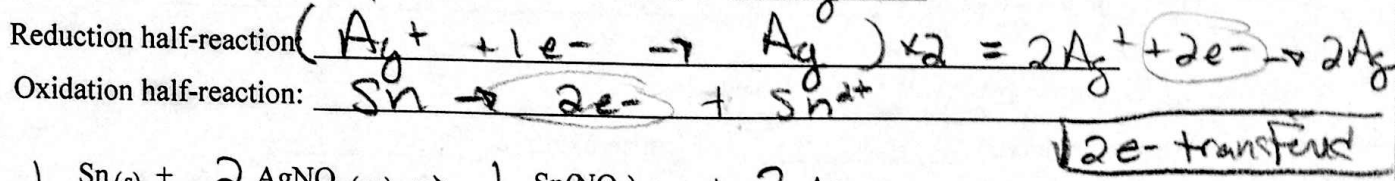
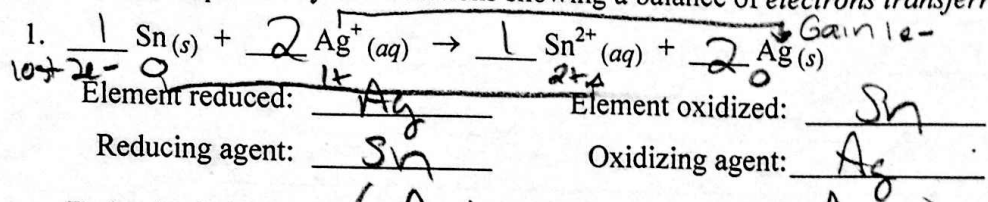
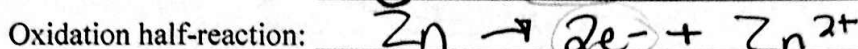
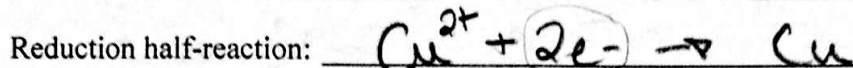
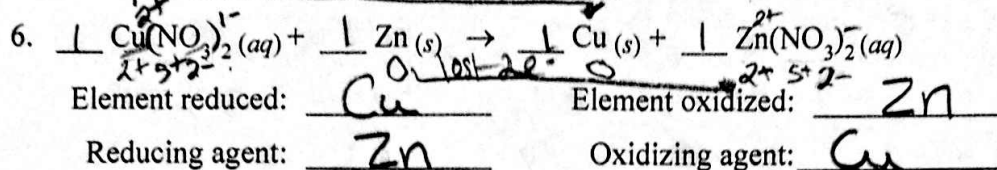
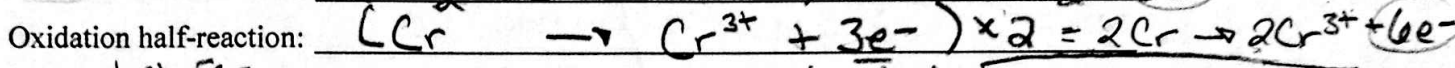
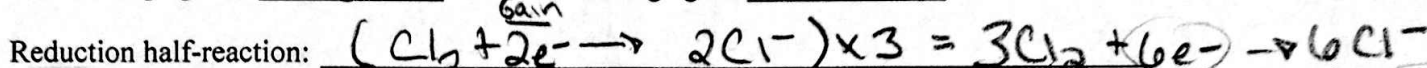
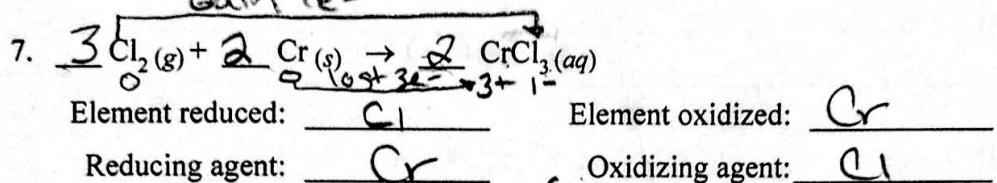


For each reaction, identify the element that is *reduced*, the element that is *oxidized*, the *reducing agent*, the *oxidizing agent*, the *reduction half-reaction*, the *oxidation half-reaction*, and then *balance* the chemical equation by half-reactions showing a balance of *electrons transferred*. *Show all work*.

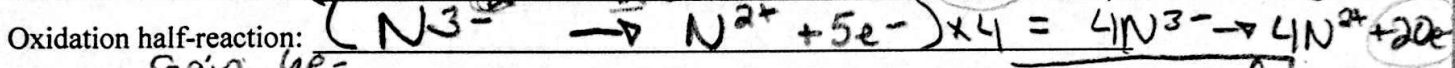
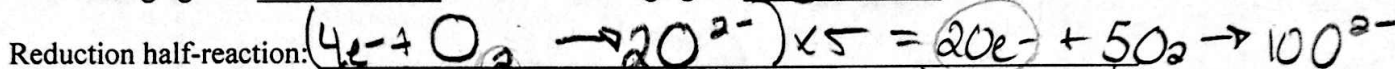
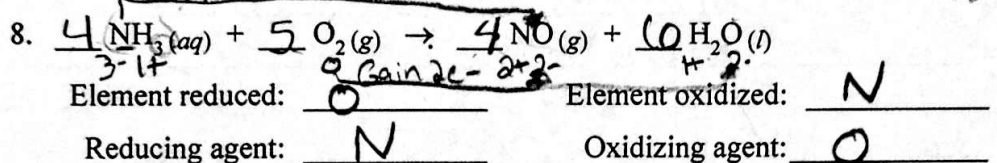




2e- transferred

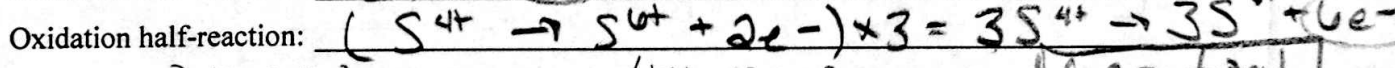
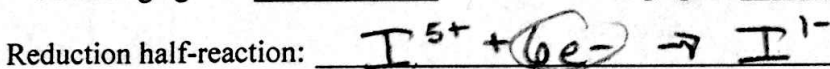
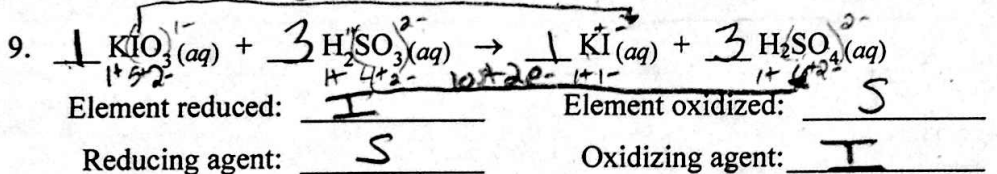


6e- total exchanged

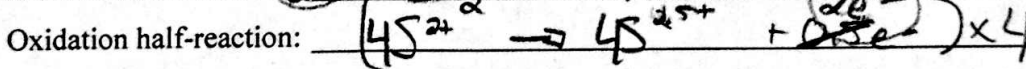
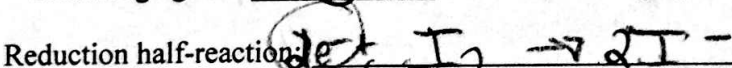
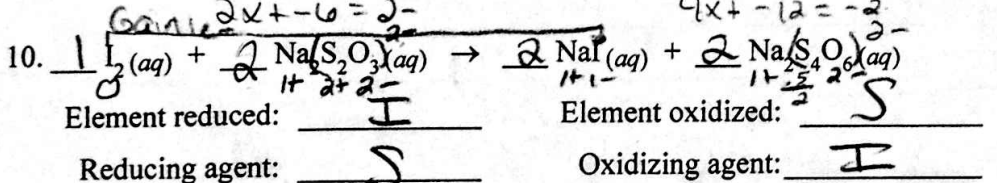


Gain 6e-

20e- total exchanged



6e- total



2e- transferred