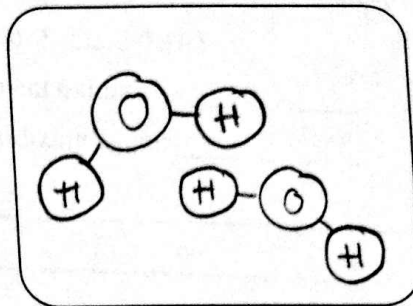
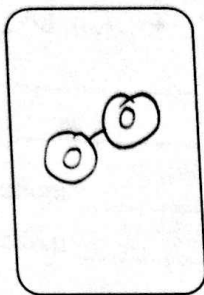
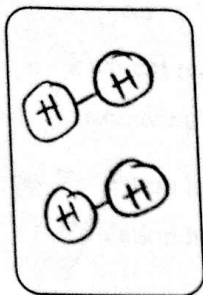


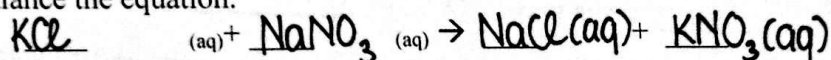
Section 8.2: Team Learning Worksheet

1. Why is the formation of water evidence of a chemical reaction? Use molecular-level sketches in your explanation.

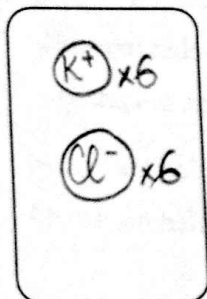


2. Is mixing aqueous potassium chloride with aqueous sodium nitrate a chemical reaction? Explain your answer.

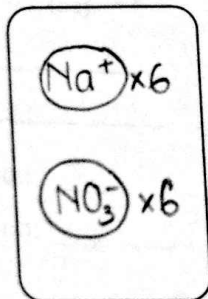
a. Balance the equation.



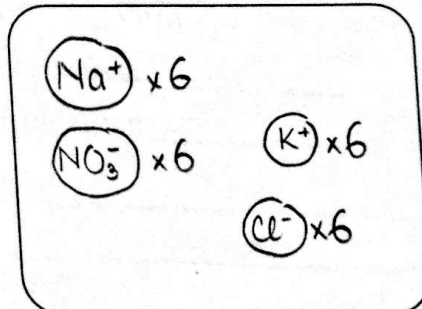
b. Show what ions are present. Remember that ionic compounds in the aqueous phase have the ions separated from each other. In the solid phase the ions form a crystal lattice. You do not need to show the water molecules. Draw 6 of each ion.



+



→ □



3. Give an example of an acid-base reaction that forms an insoluble solid. Use molecular-level sketches in your explanation. Write the molecular, complete ionic, and net ionic equations. $\text{Ca}(\text{OH})_2(\text{s}) + \text{H}_2\text{CO}_3(\text{aq}) \rightarrow 2\text{H}_2\text{O}(\text{l}) + \text{CaCO}_3(\text{s})$

4. What is the net ionic reaction of a strong acid reacting with a strong base?

