**Bond Types and Electronegativity Worksheet Name:**

**Honors Chemistry Period: \_\_\_\_\_\_ Date:\_\_\_\_\_\_\_\_\_\_**

*Read “The Bare Essentials of Polarity” comic strip. Then, use the information you learned, as well as what you remember from chapter 11 and term 1, to answer the questions below.*

1. Explain what is meant by the term “chemical bond.”
2. What are the three types of bonds represented on the third page of the comic strip? What happens to the bonding electrons in each type of bond?
3. Explain why there are four scoops of ice cream in the illustration of O2 on the third page.
4. Explain the term electronegativity in your own words. What are the trends across and down the periodic table for electronegativity?
5. Circle the atom in each pair below that will attract shared electrons more strongly.

a. C or Cl b. Rb or Br c. I or In d. Ag or S e. As or Na f. H or Se

1. What does electronegativity have to do with polarity?
2. Explain how you can use the periodic table to predict relative polarity of bonds. For example, how do you know by looking at the periodic table that a C—O bond is more polar than an N—O bond?
3. Arrange these bonded pairs in order of increasing polarity from the least polar to the most polar:

C—H C—S H—F C—N C—O H—Br

1. Name 3 things that the picture of CO2 on the fourth page illustrates about the molecule?
2. What does the crossed arrow represent in the comic strip?
3. What are 2 definitions of dipole given in the comic strip?