For each reaction,

1. Predict the possible products for each “double replacement” reaction.
2. Use the solubility rules to predict the phase of each product. Write the phase for each product.
3. Balance the complete chemical equation.
4. Write the net ionic equation below the complete chemical equation.
5. \_\_\_\_NaCl (aq)+ \_\_\_\_AgNO3 (aq) 🡪

net ionic equation:

1. \_\_\_\_KOH (aq) + \_\_\_\_Pb(NO3) 2 (aq) 🡪

net ionic equation:

1. \_\_\_\_CaCl2 (aq) + \_\_\_\_NaOH (aq) 🡪

net ionic equation:

1. \_\_\_\_Fe(NO3) 3 (aq) + \_\_\_\_Na2S (aq) 🡪

net ionic equation:

1. \_\_\_\_K2SO4 (aq) + \_\_\_\_BaCl2 (aq)🡪

net ionic equation:

1. \_\_\_\_Li2CO3 (aq) + \_\_\_\_Pb(NO3) 2 (aq) 🡪

net ionic equation:

1. \_\_\_\_MgF2 (aq) + \_\_\_\_KNO3 (aq) 🡪

net ionic equation:

1. \_\_\_\_K3PO4 (aq) + \_\_\_\_CaBr2 (aq) 🡪

net ionic equation:

1. \_\_\_\_ (NH4) 2CO3 (aq) + \_\_\_\_NaCl (aq)🡪

net ionic equation:

1. \_\_\_\_HCl (aq)+ \_\_\_\_Pb(NO3) 2 (aq) 🡪

net ionic equation: