	I	Block	Date		
Review for Unit 2 T	Test - Ch 15.1, 4 and 7				
Daigneault Chemis	try				
Station 1 – GO OV	ER ALL WORK FRO	M THE UNIT	e kilologija esemb		
1. Go through all	of the work listed on the	study guide			
a. Fill in any	portions you did not get t	o or finish			
b. Check all	answers on the keys				
c. Mark with	a star questions you find	helpful or woul	d like to go back	to and use for revi	ew.
Station 2 - GO OV	ER YOUR STUDY GI	IIDE/ REVIS	F.		
1. Go over your s	study guide with the mem	bers of your gro	oun		
a. Did yo	ou include all of the follow	ving sections:	up.		
i.	Vocabulary section				
	Include all bold vocabular	y from each secti	on that you need to	review as well as ar	ny words you
	not understand from the u	nit.			
	See the end of the chapter	Toview pages for	a word bank.		
ii.	o willie Hotes section				
	From each section we cov	ered write bullete	d notes.		
	Use the end of the chapter	review as guiden	ines for ideas on w	hat to include.	
iii.					
	Redo at least 2 examples f	rom each of the in	n class and homew	ork problems assigne	ed. Place then
	at the end of your study gt	iide or you may re	eprint the workshee	ets and attach them to	o your study
	guide.				
Belief was Direction	Complete 1-3 of the blue s still need to review and ch	section review que	estions at the end of	f the chapter for the	sections you
h		eck the answers.			sections you
U. Mark	additions and or edits to y	our study guide	in the copy of the		sections you
c. Mark	additions and or edits to y all areas you would like to	our study guide	in the copy of the		seemons you
c. Make c. Mark Station 3 – LABS 1. Candle Lab React	additions and or edits to y all areas you would like to	our study guide	in the copy of the		sources you
c. Mark Station 3 – LABS 1. Candle Lab React	additions and or edits to y all areas you would like to	eck the answers. Our study guide o go back and st	in the copy of the		socions you
Station 3 – LABS 1. Candle Lab React $C_{25}H_{52}(s) + 38 O_2(g)$ Res	additions and or edits to y all areas you would like to ion → 26 H ₂ O (g) + 25 CO ₂ (gaction	eck the answers. Our study guide o go back and st	udy for your test.	Service Services	
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2. Crystal Lab Reaction			
Reaction Balance the Reaction	Chemical or Physical Change	Reactant(s) Use the name	Product(s) Use the name
$KAI(SO_4)_2 \cdot 12H_2O(s) \rightarrow K^+(aq) + AI^{3+}$ $(aq) + SO_4^{2-}(aq) + 12H_2O(I)$	physical	alum(s)	potassium ion (aq) aluminum ion (aq) wat sulphate ion (aq)
harraman			to a compared at the state of t
Describe the state of each reactant and product sold. Discuss the ratio between each of the reactants. Draw what is happening	→ agu	+ 6	molecular compound

Describe using a drawing how your crystal is growing in your cup. Be very specific. Show all ions.

disaus with your partner(s)

answer will vary.

Describe two conditions you are using to grow your crystal and how these conditions are affecting its growth. Why are you using these conditions? What is your goal?

Use at least 5 of the following words in your description. saturated, unsaturated, temperature, evaporation, heat, clear, cloudy, shape, concentration, solute, solvent

Condition 1: raise the temperature of water and add the alum to make a saturated solution (solution with very high concentration of solute)

Condition 2: allow the solution to caol do so so that the solute (alum) can precipital out of the solution. The solvent will also prevaporate to speed up this process.

to make seed crystal

Condition 3: the position of the seed crystal in the cup will affect the shape of the crystal (the ideal shape is octahedral)

Condition 4: the rate of austal forming affect the clarity of aystal. Slow-dearer

Station 4 - BALANCING AND DRAWING CHEMICAL REACTIONS

1. Complete the information below:

Reaction Balance the Reaction	Chemical or Physical Change	Reactant(s) Use the name	Product(s) Use the name
\underline{A} Fe (s) + $\underline{3}$ O ₂ (g) \Rightarrow $\underline{2}$ Fe ₂ O ₃ (s) Rust©	chem	iron(s) oxygen(g)	tren(II) oxide

Describe what is happening by:

Classify the reactant or product as an atomic element, molecular element, ionic compound or a molecular

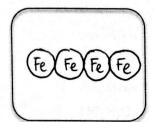
compound
atomic rement

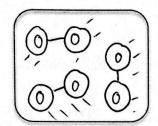
Describe the state of each reactant and product

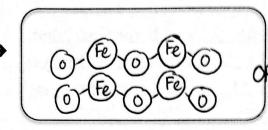
Solid

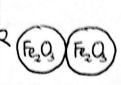
Discuss the ratio between each of the reactants and product

Draw what is happening





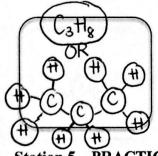


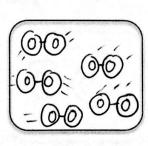


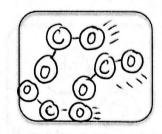
2. Complete the information below:

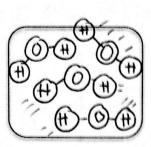
Reaction Balance the Reaction	Chemical or Physical Change	Reactant(s) Use the name	Product(s) Use the name
$1 \text{ CoH}_{\bullet}(g) + 5 \text{ Or}(g) \Rightarrow 3 \text{ COr}(g) +$		propone(g)	carbon dioxide (
$\frac{\int C_3H_8(g) + \int O_2(g) \rightarrow \int CO_2(g) + \int H_2O(g)}{4 H_2O(g)}$ your propane gas grill	chem	oxygen(g)	water(g)

Draw what is happening









Station 5 – PRACTICING NOMENCLATURE

1. Write the formula and indicate the type of bonding found in the following:

		Formula	Classification	Type of bonding
a.	sodium fluoride	NaF	ionic compound	ionic
b.	dinitriogen petoxide	N ₂ O ₅	molecular	molecular covalent
c.	lead	Pb	atomic element	metallic
d.	diamond	_ C	molecular element	network covalent

Name			Block	Date	
	e.	cupric phosphate	Cu (PO)	ionic	_ionic_
	f.	steel	Fo.	atomic element	_metallic_
	g.	ammonium perchlorate	NH CLO.	ionic	ionic
	h.	vanadium (V) fluoride	VF-	ioniccompound	ionic
	i.	calcium chloride	Cacl	ionic compound	ionic
	j.	oxygen	();	molecular	molecular covalent
2.	W	rite the name of the followin Name	g and identify the ty		Type of bonding
	a.	KMnO4 potassium	ximanganate	ionic compound	<u>ionic</u>
	b.			acid	
	c.	PCl ₃ phosphorus	trichloride	molecular compound	molecular covalent
	d.	CsClO4 Cesium per		ionic compound	ionic_
	e.	FeBr ₃ iron (III)	bromide	ionic compound	ionic_
	f.	HCN hydrocyar	nic acid	acid	
	g.	Co(NO ₃) ₃ Cabalt (I	I) nitrate	ionic compound	<u>ionic</u>
	h.	CBr4 carbon t	etra bromide	molecular compound	molecular covalent
	i.	Al ₂ (SO ₄) ₃ aluminu	m sulphate	ionic compound	ionic
	j.	HC2H3O2 acetic ac	aid	acid	<u> </u>
		on 6 – PUT IT ALL TOG rite and balance an equation for Aqueous barium chloride a aqueous solution of sodium	or the following. nd aqueous sodium:	sulfate react to form solid bari	um sulfate and an
		Balle(aq) +	Na2504(aq	$) \longrightarrow BaSO_4(as)$) + 2 Na Claq)
	b.	Solid magnesium and oxyg	en gas react in the p	resence of heat to form magne	sium oxide.
	D	2 Mg(s) + O	$(g) \xrightarrow{\Delta}$	2 Mg, O(s)	
	ч ₉)	(Mg)	9)	(Mg)-(O) (Mg)-(O)	
2.	Ba a.	lance the following equation a 1 Fe ₂ O ₃ (s) + 2	and rewrite in a com Al (s) → Al ₂ 0	plete sentence. $O_3(s) + \underline{2}$ Fe (1)	ermite!
solid v	ton	(II) oxide reacts with	h solid alumin	rum to form solid al	uminum oxide and liquid
				3_ NaCl (aq) + CrPO	
aqueoris	80	dum phosphate react	s with chromiu		oform aqueous radium nd chromium phosphatesolic