poves for the reaction below.

Review for Unit 3 Test - Ch 8 Daigneault Chemistry

station 1 - GO OVER ALL WORK FROM THE UNIT

- 1. Go through all of the work listed on the study guide.
 - a. Fill in any portions you did not get to or finish.
 - b. Check all answers on the keys.
 - c. Mark with a star questions you find helpful or would like to go back to and use for review.

Station 2 - GO OVER YOUR STUDY GUIDE/ REVISE

- 1. Go over your study guide with the members of your group.
 - a. Did you include all of the following sections:

i. Vocabulary section

Include all bold vocabulary from each section that you need to review as well as any words you do not understand from the unit. See the end of the chapter review pages for a word bank.

ii. Outline notes section

From each section we covered write bulleted notes. Use the end of the chapter review as guidelines for ideas on what to include.

iii. Example section

Redo at least 2 examples from each of the in class and homework problems assigned. Place them at the end of your study guide or you may reprint the worksheets and attach them to your study guide.

Complete 1-3 of the blue section review questions at the end of the chapter for the sections you still need to review and check the answers.

dditions and or edits to your study guide.

b. Make additions and of edits to your steady gardeness. c. Mark all areas you would like to go back and study for your test.
Station 3 - Precipitation and Acid Base Reactions 1. Randomly pick one cation and one anion from the pile. Write the formula
4. Complete the following for the reaction above and draw in the boxes below.
a. Write out the balanced equation. Ph(N)_2/2(a)+21 Class -> PhClass +21 NO3(a) a. Write out the ionic equation. Ph'(ag) + 2NO3(ag) + 2Class -> PhClass + 2Liteap + 2NO3(ag) b. Write out the net ionic equation.
c. What are the spectator ions from the reaction?
Litand NO3
(i+ C1-) (i+ C1-) (i+ Ci+ NO3-NO3-+(Eg))

sulfuric acid reacts with potassium hydroxide to form salt water. d. Write out the balanced equation. On Brattotle) + h2804 (ag) e. Write out the ionic equation. 20H(ag) + 2H2O(e) + 2Kt cags + SOyd ca f. Write out the net ionic equation. g. What are the spectator ions from the reaction? Station 4 - Correct my mistakes One of the best ways to learn is to correct someone else's mistakes. Correct my mistakes below. 1. Complete the products and identifying reaction types. 2LiCl $(s) \rightarrow 2$ 3 CuCO₃ (aq) + 2 Al (s) \rightarrow _ * Balance * should replace mitals -> reaction c. $Mg(s) + \underline{\hspace{1cm}} O_2(g) \rightarrow \underline{\hspace{1cm}}$ after corrected (3) $C_2H_5OH(I) +$ *Not balanced correctly, there $HF(aq) + Ca(OH)_2(aq) \rightarrow (24)$ are 11-0 on the left. Bonus: Who is the celebrity chemical in Ch 8? Why? to balance the charges correct Ca CO2 - Be cause the dissolution and reaction formula of the limistone as it goes through the Soil Forms staladition Station 5 - RedOx - Who? - What? - How?

complete the following and draw in the boxes for the reaction below.

Station 5 - RedOx - Who? - How?- What?
1. In the following redox reaction identify: Who is involved in the redox. What is the
hamber of electrons that are transferred and How are the electrons are transferred using
half reactions. The second and now are the electrons are transferred using the second and now are the electrons are transferred using the second and now are the electrons are transferred using the second and now are the electrons are transferred using the second and now are the electrons are transferred using the second and now are the electrons are transferred using the second and now are the electrons are transferred using the second and now are the electrons are transferred using the second and now are the electrons are transferred using the second and now are the electrons are transferred using the second and now are the electrons are transferred using the second and now are the electrons are transferred using the second and now are the electrons are transferred using the second and now are the electrons are transferred using the second and now are the electrons are transferred using the second and now are the electrons are transferred using the second and now are the electrons are transferred using the second and transferred using the second and the second and the second and the second are transferred using the second and the second are transferred using the second and the second and the second are transferred using the second are transferred using the second and the second are transferred using the second are tr
Δ Zn(s) + Δ HNO ₃ (ag) \rightarrow H ₂ (g) + Zn(NO ₃) ₂ (ag) \rightarrow + \rightarrow Q agin e -
Element reduced: H GER Element oxidized: Zn LEO
Element reduced: H GER Element oxidized: Zn LEO Oxidizing agent: Reducing agent: H
Reduction half-reaction: 2H'+ + 2e> Ha(g)
Reduction half-reaction: 2H + Le> Ha(g)
Oxidation half-reaction: $2n \rightarrow 2e + 2n^{2+}$
Oxidation nail-reaction:
Spectator Ions:
WHAT
What is the total number of electrons transferred?
Draw out the RedOx reaction below:
7909
0000 + 0 80 > 200 hydrosen gas will bubble out of
to so the state of
bubble out of
the salution
What is the driving force of this reaction?
What is the driving force of this reaction?
2. Are precipitation reactions generally oxidation-reduction reactions? Why or why not?
and a base something all and a base of a base of a last
100' and are were men and charles
No, they are not exchanging electrons and changing oxidation steele-they are combining to form in saluble 3. Single-displacement reactions are also RedOx reactions.
3. Single-displacement reactions are also RedOx reactions. Why must this be true? Explain using an example, explanation and drawing.
why must this be true? Explain using an example, explanation and drawing.
Because you are taking a cuments in their natural
State and combining them to form a compound where
when will other transfer e- or share e- as when hand
they will either transfer e- or share e- as they bend
2 Na(1) + Class - 2 Nacl (1)
(Nath CI-
Natio ocit
Na Cla

Station 6 - Putting it all together

The following examples are from your reaction types worksheets. Complete the extensions for each example.

Write letter code for the type of reaction in the first space, then complete and balance the chemical equation:

type
1. \Re : 2 Li(s) + $\mathop{\perp}$ MgBr ₂ (aq) \rightarrow 2 Li Br(ag) + Mges) Write out the net ionic equation.
Write out the net ionic equation
What are the spectator ions from the reaction? Spectator What are the spectator ions from the reaction?
What are the spectator ions from the reaction?
Br-
type
2. Record aluminum chloride reacts with silver nitrate → Silver chloride and Write out the balanced equation.
Al Cl3cags + 3 Ag NO sags - Al (NO3) 3 (eg) + 3 Ag Cl cs) Draw the reaction below:
Alst Nos Ast Nos in solution water solution water solution
CI- (AT NOS-) (ACD) ACD & Solid,
watersolution watersolution precipità
type 3.5 untrusis: 4 Na(s) + 10 0 0 0 0 0 0 0 0 0
3. Sunthisis: \subseteq Na(s) + \bigcirc O ₂ (g) \rightarrow \bigcirc Clock O(s) \bigcirc O \bigcirc Gain.
Element reduced: Oxidizing agent: Reducing agent: Reducing agent:
Reduction half-reaction: Og 2 2002-
Oxidation half-reaction: $(Na - 71e - + Na^{+1})xa\sqrt{(1a + 4a)}$
Draw out the RedOx reaction below:
TOTAL
(Cara + (Cara) + (
Solid Non goslous O2 Solid Naci
type A R
4. ☐ : acetic acid and potassium hydroxide → Worter and Potassium Draw out the Acid/ Base reaction below:
That out the Actu/ Base reaction below:
+ + + + HOTH Wester
CaHO3- OH- K+ CaHO2-